

(19) World Intellectual Property Organization International Bureau



(43) International Publication Date
6 May 2005 (06.05.2005)

PCT

(10) International Publication Number
WO 2005/039763 A1

(51) International Patent Classification⁷: B01J 31/12. (74) Common Representative: BASELL POLYOLEFINE GMBH; Intellectual Property, Industriepark Hoechst, Bldg B 413, 65926 Frankfurt am Main (DE).

(21) International Application Number: PCT/EP2004/011316 (81) Designated States (unless otherwise indicated, for every kind of national protection available): AE, AG, AL, AM, AT, AU, AZ, BA, BB, BG, BR, BW, BY, CZ, DE, DK, DM, DZ, EC, EE, EG, ES, FI, GB, GD, GE, GH, GM, HR, HU, ID, IL, IN, IS, JP, KE, KG, KP, KR, KZ, LK, LR, LS, IT, LU, IV, MA, MD, MG, MK, MN, MW, MX, MZ, NA, NI, NO, NZ, OM, PG, PH, PL, PT, RO, RU, SC, SD, SE, SG, SK, SL, SY, TJ, TM, TN, TR, TT, TZ, UA, US, UZ, VC, VN, YU, ZA, ZM, ZW.

(22) International Filing Date: 9 October 2004 (09.10.2004)

(25) Filing Language: English

(26) Publication Language: English

(30) Priority Data: 103 48 624.0 15 October 2003 (15.10.2003) DE
60/518,427 6 November 2003 (06.11.2003) US

(71) Applicant (for all designated States except US): BASELL POLYOLEFINE GMBH [DE/DE]; Brühler Strasse 60, 50389 Wesseling (DE).

(72) Inventors; and

(75) Inventors/Applicants (for US only): MIHAN, Shahram [IR/DE]; Am Eichwald 18, 65812 Bad Soden (DE); ROSENDORFER, Philipp [DE/DE]; Dochmahlstrasse 30, 67434 Neustadt (DE); KARER, Rainer [DE/DE]; Oskar-Schlemmer-Ring 48, 67657 Kaiserslautern (DE); SCHNEIDER, Martin [DE/DE]; Hornauer Str. 37b, 65779 Kelkheim (DE); ECK, Peter [DE/DE]; Brunnen-gasse 4, 67098 Bad Dürkheim (DE).

(84) Designated States (unless otherwise indicated, for every kind of regional protection available): ARIPO (BW, GH, GM, KE, LS, MW, MZ, NA, SD, SL, SZ, TG, UG, ZM, ZW), Eurasian (AM, AZ, BY, KG, KZ, MD, RU, TJ, TM), European (AT, BE, BG, CH, CY, CZ, DE, DK, EE, ES, FI, FR, GB, GR, HU, IE, IT, LU, MC, NL, PL, PT, RO, SE, SI, SK, TR), OAPI (BF, BJ, CF, CG, CI, CM, GA, GN, GQ, GW, ML, MR, NE, SN, TD, TG).

Declaration under Rule 4.17:

— of inventorship (Rule 4.17(iv)) for US only

Published:

— with international search report

For two letter codes and other abbreviations, refer to the "Guidance Notes on Codes and Abbreviations" appearing at the beginning of each regular issue of the PCT Gazette.

A1

WO 2005/039763 A1

(54) Title: SUPPORTED METAL ALKYL COMPOUND AND ITS PREPARATION

(57) Abstract: Pulverulent solid which consists essentially of at least one metal alkyl compound bound chemically and/or physically to a finely divided, porous, mechanically stable and chemically inert support, has a proportion by weight of metal alkyl compound of at least 5% by weight, based on the support, and has an angle of repose, determined in accordance with ISO 4324, of up to 48°. The solid allows trouble-free metering as active component into a reactor.

Supported metal alkyl compound and its preparation

5 The present invention relates to a pulverulent solid consisting essentially of at least one metal alkyl compound which is chemically and/or physically bound to a finely divided, porous, mechanically stable and chemically inert support. It further provides a process for preparing this solid and to its use in a polymerization process.

10 Solids of this type are known, for example, from the Japanese published specification JP-A-53-011 982 or from DE-A-26 23 693.

The solid described in JP-A-53-011 982 is magnesium oxide, zinc oxide, silica gel, magnesium hydroxide, aluminum silicate or magnesium silicate which has been reacted with triethylaluminum, 15 tri-n-octylaluminum, diethylaluminum hydride or isobutylaluminum dihydride at from 50 to 100°C in hexane or decane. The resulting solids or supported aluminum compounds are used as activators for Phillips catalysts comprising chromium oxide in ethylene polymerization. For the aluminum alkyl compounds to be able to display their known activation effect, they have to react with the chromium oxide of the Phillips catalysts. It is therefore a physical requirement that the solid of 20 JP-A-53-011 982 contains only a very small amount, if any, of immobilized aluminum alkyls, since otherwise the mass transfer between the solid and the Phillips catalyst which is essential for activation would no longer take place. Nothing is said in JP-A-53-011 982 about a purifying effect of the solid.

25 The solid described in DE-A-26 23 693 is likewise based on supports such as alumina, hydrated alumina, silicic acid, aluminum silicate and magnesium silicate and, in addition, calcium carbonate, magnesium carbonate or a polyolefin powder which have been impregnated with at least one liquid aluminum alkyl compound having a vapor pressure of less than 1 mm/80°C, e.g. tri-n-octylaluminum. These solids are used as cocatalysts for the (co)polymerization of 30 α -monoolefins in the gas phase by means of Ziegler or Phillips catalysts comprising transition metals.

EP-A-560128 discloses a solid which contains metal alkyl groups and in which the metal alkyls are all bound to the surface of the support. Although this has the advantage that it reduces the 35 interactions with the polymerization catalyst, the supported metal alkyl compounds obtained in this way have, like all other previously known solids comprising metal alkyls, the disadvantage that they easily become conglutinated when being metered into the reaction, in particular when using long, narrow lines, and are therefore not very suitable for metering into a gas-phase fluidized-bed reactor.

It is an object of the present invention to overcome the abovementioned disadvantages of the prior art and to provide a supported metal alkyl compound which allows trouble-free metering into a reactor, in particular into a gas-phase fluidized-bed reactor.

- 5 We have found that this object is achieved by a novel solid which is based on a metal alkyl compound bound chemically and/or physically to a finely divided, porous and mechanically stable support, has a metal alkyl content of more than 5% by weight and has an angle of repose, determined in accordance with ISO 4324, of not more than 48°.
- 10 Due to its excellent powder flow properties, the solid of the present invention is suitable for metering through even narrow and/or long feed lines without blockages occurring.

In the following, this novel solid based on the metal alkyls and the supports will in the interest of brevity be referred to as supported metal alkyl.

- 15 It is important, firstly, that the supported metal alkyl of the present invention has an angle of repose determined in accordance with ISO 4324 of not more than 48°. The angle of repose is a suitable measure for the powder flow of the supported metal alkyl and thus also a measure of its suitability in respect of metering.
- 20 Secondly, the supported metal alkyl has at least 5% by weight, preferably from 10 to 40% by weight, particularly preferably from 15 to 30% by weight (based on the support), of the metal alkyl compound on the surface of the support. The solids of the present invention thus contain an amount of metal alkyl compound which is significantly above that corresponding to saturation of
- 25 the free bonding sites of the solid, so that metal alkyl which is not chemically bound is also present on the support surface. According to the present invention, the term physical bonding refers to any attractive, nonchemical interaction which is able to hold the metal alkyl on the surface of the support.
- 30 The main basis of the supported metal alkyl is formed by a monovalent, divalent or higher-valent metal alkyl. Accordingly, all metal alkyls of 1-, 2-, 3-, 4- and 5-valent metals, but in particular the alkyls of 2-, 3- and 4-valent metals, are suitable. Examples of well-suited metal alkyls are the alkyl compounds of beryllium, magnesium, calcium, strontium, barium, zinc, boron, aluminum, gallium, indium, thallium, tin and lead. Among these, the alkyl compounds of boron, aluminum and zinc are
- 35 particularly advantageous and are therefore very particularly preferably used. The alkyl compounds of aluminum in turn offer very particular advantages for the preparation of the solid of the present invention and are therefore accorded special importance and are very particularly preferably employed.

In general, the alkyl groups of the metal alkyls have from 1 to 20 carbon atoms and may also be substituted as long as the substituent is compatible with the support and the intended use. Alkyl groups having from 2 to 8 carbon atoms, e.g. ethyl, propyl, isopropyl, n-butyl, isobutyl, n-pentyl, n-hexyl, heptyl, octyl and 2-methylhexyl and also cyclohexyl and methylcyclohexyl are very

5 particularly advantageous and are therefore very particularly preferably used. Apart from fully alkylated metal alkyl compounds, it is also possible to use partially alkylated metal alkyl compounds, with possible further substituents being, for example, halogens, in particular chlorine and iodine, hydrides or alkoxy groups.

10 Examples of metal alkyls which are preferably used for preparing the solid of the present invention are trimethylaluminum, triethylaluminum, tripropylaluminum tributylaluminum, triisobutylaluminum, trihexylaluminum, trioctylaluminum, tridodecylaluminum diethylaluminum chloride, diethylaluminum hydride, ethylaluminum sesquichloride ethylaluminum dichloride, diisobutylaluminum chloride, isobutylaluminum dichloride, diethylaluminum iodide,

15 dipropylaluminum chloride, diisobutylaluminum hydride, diethylaluminum ethoxide, dimethylaluminum chloride, methylaluminum sesquichloride, diethylzinc, triethylboron, butyloctylmagnesium and also partially hydrolyzed metal alkyls such as methylaluminoxane(MAO) or isobutylaluminoxane (IBAO). Particular preference is given to triethylaluminum, trisobutylaluminum and trihexylaluminum, diethylzinc and diethylaluminum ethoxide. The metal alkyls are compounds which are known per se and are commercially available.

20 The further important basis of the supported metal alkyl is formed by the finely divided, porous and mechanically stable support. This support or the particles of which it consists is/are abrasion resistant, which is very important for use of the solid of the present invention in a fluidized or stirred gas phase or in a suspension with turbulent flow, and the support is also not damaged under its own weight during storage.

The support to be employed according to the present invention preferably bears, in the unladen state, i.e. in the state free of metal alkyl, functional groups on its surface which can effect

25 chemical and/or physical bonding between the metal atoms of the metal alkyl and the surface of the support. Furthermore, it is important for the solid of the present invention that only the surface of the support reacts with the metal alkyls and the remainder of the support is chemically inert toward these.

30 Possible supports include both polymeric organic solids and polymeric and/or crystalline inorganic solids.

Inorganic oxides suitable as supports may be found among the oxides of elements of groups 2, 3, 4, 5, 13, 14, 15 and 16 of the Periodic Table of the Elements. Preference is given to oxides or

mixed oxides of the elements calcium, aluminum, silicon, magnesium, titanium, zirconium or hafnium and also corresponding oxide mixtures. Other inorganic oxides which can be used either alone or in combination with the abovementioned oxidic supports are, for example, ZrO_2 or B_2O_3 . Preferred oxides are silicon dioxide, in particular in the form of a silica gel or a pyrogenic silica, or 5 aluminum oxide. A preferred mixed oxide is, for example, calcined hydrotalcite. Further inorganic support materials which may be mentioned are, for example, magnesium chloride, aluminum oxide hydrate, aluminum phosphate or metal silicates.

In general, inorganic supports which, like silicates or aluminosilicates, have a porous structure 10 with a large surface area have a mean particle size of from 1 μm to 1 mm, preferably from 10 to 100 μm , in particular from 20 to 70 μm . The support advantageously has a BET surface area of from 10 to 1000 m^2/g , preferably from 50 to 500 m^2/g and in particular from 200 to 400 m^2/g . It is also advantageous for its pore volume to be from 0.1 to 5 ml/g , preferably from 0.8 to 3.5 ml/g , particularly preferably from 0.9 to 2.5 ml/g . Exceptions are supports whose surfaces are formed 15 only by means of swelling, for example in the case of sheet silicates or hydrotalcites.

The inorganic support can be subjected to a thermal treatment, for example to remove adsorbed water. Such a drying treatment is in general carried out at from 80 to 300°C, preferably from 100 to 200°C, preferably under reduced pressure and/or in a stream of inert gas, for example nitrogen 20 or argon. The inorganic support can also be calcined, in which case the concentration of OH groups on the surface can then be set and the structure of the solid can possibly be altered by treatment at from 200 to 1 000°C.

The inorganic support material can also be modified chemically. For example, treatment of silica 25 gel with NH_4SiF_6 leads to fluorination of the silica gel surface or treatment of silica gels with silanes containing nitrogen-, fluorine- or sulfur-containing groups gives correspondingly modified silica gel surfaces. Further treatment methods are described in WO 00/31090.

Examples of organic solids which can be used as supports are modified and unmodified cellulose 30 and starch, modified and unmodified lignin, synthetic polymers which contain carboxylic acid, sulfonic acid, phosphonic acid, acid phosphoric ester, hydroxy and/or primary and/or secondary amino groups and are based on polyethylene and ethylene copolymers, polypropylene and propylene copolymers, vinyl chloride copolymers, polystyrene and styrene copolymers, polyphosphazenes, polyamides, polycarbonates, polyalkylene terephthalates, polyphenylene 35 oxides, polysulfones, polyphenylene sulfides, polyaryl sulfones, polyether sulfones, polyaryl ethers, phenol-formaldehyde resins (phenolic resins), urea-formaldehyde resins (amino resins), melamine-formaldehyde resins, melamine-phenol-formaldehyde resins, silicones, polyimides, epoxy resins or crosslinked polyurethanes. It is also possible to use functionalized polymeric supports, e.g. supports based on polystyrenes. Preferred organic support materials are finely 40

divided polymer powders, for example powders composed of polyolefins such as polyethylene or polypropylene or of polystyrene.

According to the present invention, the supported metal alkyl consists essentially of the two
5 above-mentioned components metal alkyl compound and support, with small amounts of additional components also being able to be present. Such additional components can be additives or auxiliaries of any type, for example antistatics.

To obtain the particularly good powder flow properties of the supported metal alkyl, it is
10 advantageously prepared from the metal alkyls and supports described in detail above using the process of the present invention. According to this, the pulverulent solid is obtainable by firstly drying the support to a water content below 3% by weight, preferably 2% by weight, particularly preferably 1% by weight, if the original water content is above this, and then bringing the metal alkyl compound into contact with the support in a solvent having a boiling point which is not above
15 30°C, followed by removal of the solvent so that the supported alkyl compound remains and a particularly free-flowing pulverulent solid is obtained.

For the present purposes, the water content is the total content of hydroxyl groups, expressed as water, which are able to react with the metal alkyl. The water content can be determined, for
20 example, by titration with the metal alkyl itself.

As solvent to be used in the process of the present invention, it is possible to employ any liquid which is inert toward the metal alkyl compound and has a boiling point less than or equal to 30°C. Preference is given to solvents which are inert toward metal alkyl compounds and have a boiling
25 point in the range from 10 to 30°C. Particular preference is given to isopentane, which has a boiling point of 28°C.

Particularly good results are achieved when the support is firstly suspended in the solvent, in particular isopentane, and the metal alkyl compound is subsequently added in diluted form. As an alternative, the metal alkyl compound can be dissolved beforehand in the isopentane.
30

The solids of the present invention, in particular those obtained by the process of the present invention, are mechanically stable and are very suitable for the purification of gaseous substances and also of chemical reactors in which air- and/or water-sensitive substances are reacted with one
35 another. They have the advantage over other purifying agents that they eliminate a wide variety of contaminants even in the ppm range, after which they can easily be separated off again from the gaseous substances and be taken from the reactors again.

Substances which are particularly suitable for purification by the solids of the present invention include all inert gases, hydrogen and also polymerizable monomers containing ethylenically unsaturated groups. The reactors are, in particular, plants for the (co)polymerization of such monomers. Here, the solids of the present invention are especially suitable for removing 5 impurities in plants which are intended for the (co)polymerization of olefins in the fluidized or stirred gas phase and contain polymer particles being mixed.

The present invention further provides a process for preparing homopolymers and copolymers of α -olefins in a gas-phase fluidized-bed reactor, in which the α -olefin is (co)polymerized in a 10 polymerization zone of the gas-phase fluidized-bed reactor at from 30 to 125°C and pressures of from 1 to 100 bar in the gas phase in a mixed bed of finely divided polymer in the presence of a catalyst comprising a transition metal and using the pulverulent solid of the present invention and discharging the resulting (co)polymers from the reactor.

15 In this process for preparing homopolymers and copolymers of α -olefins, the monomer or monomers concerned is/are fed continuously or discontinuously into the reactor in a known manner. For this purpose, it is possible to use all customary and known catalysts, in particular Phillips, Ziegler and single-site catalysts such as metallocene catalysts as are described, for example, in the prior art discussed at the outset and in US-A-4 532 311 or US-A-3 709 853.

20 Among metallocene catalysts, supported chromium(II) compounds such as bis(cyclopentadienyl)chromium(II) are particularly advantageous because they can be used directly for the (co)polymerization without an additional activation step and without addition of cocatalysts. Accordingly, the (co)polymerization can be carried out either in the presence or absence of a cocatalyst. Possible cocatalysts are essentially all the abovementioned metal alkyls, 25 but in particular the alkyl compounds of aluminum, and also the supported cocatalysts mentioned at the outset. In addition, a finely divided, inert solid which is free of metal alkyls can also be present in the (co)polymerization. Further additives which can, if required, be metered into the reactor include molecular weight regulators such as hydrogen and antistatics (cf. US-A-4 532 311). The (co)polymerization can be carried out in the fluidized or stirred gas phase in 30 a mixed bed of finely divided (co)polymer. Depending on whether this process is carried out continuously or batchwise, the resulting (co)polymers are discharged continuously from the reactor or taken from the reactor after the reaction has ended.

Furthermore, the metal alkyl supported according to the present invention can also be used, in 35 particular for the removal of impurities, in a multizone reactor as is described, for example, in WO 97/04015 or WO 00/02929.

To meter the supported metal alkyls of the present invention into a gas-phase fluidized-bed reactor for the polymerization of ethylene, it is possible to use, for example, the discontinuous metering apparatus described in EP-A-226935.

5 The solid of the present invention can be introduced into the plants or reactors before carrying out these processes in order to free the reactor of oxygen, water and other interfering compounds. Here, the solid of the present invention displays its excellent purifying action. The solid of the invention can, however, also be introduced during the (co)polymerization to influence the time-activity behavior. For the purposes of the present invention, the time-activity behavior is the
 10 change in activity of the catalyst used as a function of time. This can be altered by means of the supported metal alkyls of the present invention. Thus, for example, in the case of Phillips catalysts the reduction of chromium(VI) which is necessary for activation takes a relatively long time when it is brought about by means of ethylene. In contrast, when unsupported metal alkyls such as aluminum alkyls are used, the catalyst starts acting undesirably quickly. The use of supported
 15 metal alkyls, on the other hand, makes it possible to achieve a gentle but nevertheless rapid commencement of the reaction within a few minutes.

Examples

20 Various metal alkyl compounds were supported and their powder flow behavior was determined.

The method of ISO 4324 was used to determine the powder flow. ISO 4324 is a standard which is actually intended for testing the angle of repose of surfactants in powder or granule form, but can also be applied to other powder and granulated materials having comparable properties. The
 25 determination of the angle of repose was carried out under an inert atmosphere.

The powder flow of the pulverulent solid is better the smaller the angle of repose. The angles of repose were, in the interests of simplicity, divided into the following powder flow classes:

	Angle of repose [°]	Powder flow class	Assessment
30	up to 44	1	very good
	44-46	2	good
	46-48	3	slightly sticky
	48-50	4	poor
35	above 50	5	very poor

Example 1:

The silica gel (SG 332, pore volume: 1.8 ml/g, mean particle size: 50 μ m, specific surface area determined by the BET method: 325 m^2/g) was firstly heated at 250°C for 6 hours under an argon atmosphere. The water content was determined volumetrically by titration of the OH groups with triethylaluminum, giving a value of 1.4% by weight. 33.5 g of the dried silica gel were subsequently suspended in 220 ml of Isopentane under a protective gas atmosphere in a 500 ml flask.

10 10.05 g (30% by weight based on the support) of undiluted trihexylaluminum were then slowly added dropwise to this suspension. As a result of the exothermic reaction, the temperature rose from 12°C to 27°C. After the trihexylaluminum had been added, the mixture was stirred at not more than 30°C for another 1 hour and the Isopentane was removed at room temperature under reduced pressure.

15 As can be seen from table 1, a fine product of powder flow class 1 having very good powder flow was obtained.

Examples 2, 3

20 Example 1 was repeated using 50% by weight and 10% by weight of trihexylaluminum. Products of powder flow classes 2 and 1 having good to very good powder flow were likewise obtained.

Comparative example C4

25 Example 1 was repeated using 4% by weight. A supported metal alkyl compound of powder flow class 4 which displayed poor flow was obtained; this is unsuitable for use in the metering apparatus for the gas-phase fluidized-bed reactor.

30 Comparative examples C5, C6, C7

Example 1 was modified in that application to the support was carried out using heptane, hexane or pentane instead of Isopentane as solvent. Supported metal alkyl compounds of powder flow class 4 or 5 which displayed very poor flow were obtained in each case; these are unsuitable for 35 use in the metering apparatus for the gas-phase fluidized-bed reactor.

Example 8

Example 1 was modified in that the pretreatment of the support comprised a calcination at 750°C. A supported metal alkyl compound of powder flow class 1 which displayed very good flow was 5 obtained.

Example 9

Example 1 was modified in that the pretreatment of the support comprised a calcination at 800°C 10 and the proportion of metal alkyl compound was 42.9% by weight, based on the support. A supported metal alkyl compound of powder flow class 1 was obtained.

Example 10

15 Example 1 was carried out using triethylaluminum (TEA) as metal alkyl compound. A supported metal alkyl compound of powder flow class 1 was obtained.

Example 11

20 Example 1 was carried out using triethylaluminum (TEA) as metal alkyl compound. The thermal pretreatment of the support was omitted. A supported metal alkyl compound of powder flow class 2 was obtained.

Examples 12, 13, 14

25 Example 1 was repeated using diethylzinc (DEZ), triisobutylaluminum (TIBA) or diethylaluminum ethoxide (DEALOX) as metal alkyl compound. Metal alkyl compounds of powder flow class 2 were obtained.

30

35

40

Table 1

Example	Metal alkyl	Pretreatment of supports [°C]	Amount of metal alkyl [% by weight] *	Solvent	Powder flow class	
5	1	THA	250	30	isopentane	1
	2	THA	250	50	isopentane	1
	3	THA	250	10	isopentane	2
	C4	THA	250	4	isopentane	4
10	C5	THA	250	30	heptane	5
	C6	THA	250	30	hexane	5
	C7	THA	250	30	pentane	4-5
	8	THA	750	30	isopentane	1
15	9	THA	800	42.9	isopentane	1
	10	TEA	250	30	isopentane	1
	11	TEA	none	30	isopentane	2
	12	DEZ	250	30	isopentane	2
20	13	TIBA	250	30	isopentane	2
	14	DEALOX	250	30	isopentane	1

* based on the support

We claim:

1. Pulverulent solid which consists essentially of at least one metal alkyl compound bound chemically and/or physically to a finely divided, porous, mechanically stable and chemically inert support and which has a proportion by weight of metal alkyl compound of at least 5% by weight, based on the support, and an angle of repose, determined in accordance with ISO 4324, of up to 48°.
2. A pulverulent solid as claimed in claim 1, wherein the proportion by weight of the metal alkyl compound is in the range from 10 to 40% by weight.
3. A pulverulent solid as claimed in claim 1 or 2, wherein the metal alkyl compounds are selected independently from the group consisting of alkyl compounds of the elements lithium, beryllium, magnesium, calcium, strontium, barium, zinc, boron, aluminum, gallium, indium, thallium, tin and lead.
4. A pulverulent solid as claimed in any of claims 1 to 3, wherein the support is an inorganic support, in particular a silicon dioxide, aluminum oxide, or magnesium oxide support or a mixture thereof.
5. A process for preparing a pulverulent solid as claimed in any of the preceding claims, which comprises the steps
 - if necessary drying the support to a water content of less than 3% by weight,
 - bringing the metal alkyl compound into contact with the support in an inert solvent having a boiling point of less than 30°C,
 - removing the solvent from the pulverulent solid.
6. A process as claimed in claim 5, wherein the solvent is isopentane.
7. A process as claimed in claim 5 or 6, wherein the support is suspended in isopentane and the metal alkyl compound is subsequently added in undiluted form or as a solution in isopentane.
8. A process as claimed in any of claims 5 to 7, wherein the solvent is removed at from 0 to 40°C and pressures up to 10 000 Pa.

9. A process for preparing homopolymers and copolymers of α -olefins in a gas-phase fluidized-bed reactor, in which the α -olefin is -(co)polymerized in a polymerization zone of the gas-phase fluidized-bed reactor at from 30 to 125°C and pressures of from 1 to 100 bar in the gas phase in a mixed bed of finely divided polymer in the presence of at least one catalyst comprising a transition metal and in the presence of a solid and the resulting (co)polymers are discharged from the reactor, wherein a pulverulent solid as claimed in any of claims 1 to 4 is used.
10. A process as claimed in claim 9, wherein the time-activity behavior of the catalyst used is influenced by means of the solid.
11. A process as claimed in claim 9 or 10, wherein the solid is used to remove oxygen, carbon dioxide, water and/or other interfering compounds during start-up of the gas-phase fluidized-bed reactor.

15

20

25

30

35

40

INTERNATIONAL SEARCH REPORT

International Application No
PCT/EP2004/011316A. CLASSIFICATION OF SUBJECT MATTER
IPC 7 B01J31/12 B01J31/14 C08F10/00

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)
IPC 7 B01J C08F

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practical, search terms used)

EPO-Internal, PAJ, WPI Data

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	WO 95/10542 A (EXXON CHEMICAL PATENTS INC) 20 April 1995 (1995-04-20) page 5, line 11 - page 7, line 8 page 10, line 10 - line 16. ----- US 4 370 456 A (GEORGE ET AL) 25 January 1983 (1983-01-25) column 4, line 54 - column 5, line 38; example 1 -----	1-6, 8-10
X	DATABASE WPI Section Ch, Week 197444 Derwent Publications Ltd., London, GB; Class E36, AN 1974-77020V XP002316216 & SU 409 723 A (STERLITAMAK ISOPRENE RUB) 3 June 1974 (1974-06-03) abstract -----	1-6
X	-----	1, 3, 5, 6, 11

 Further documents are listed in the continuation of box C. Patent family members are listed in annex.

* Special categories of cited documents :

- *A* document defining the general state of the art which is not considered to be of particular relevance
- *E* earlier document but published on or after the International filing date
- *L* document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)
- *O* document referring to an oral disclosure, use, exhibition or other means
- *P* document published prior to the International filing date but later than the priority date claimed

T later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention

X document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone

Y document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art.

A document member of the same patent family

Date of the actual completion of the international search

3 February 2005

Date of mailing of the international search report

17/02/2005

Name and mailing address of the ISA

European Patent Office, P.O. 5818 Patenttaan 2
NL - 2280 HV Rijswijk
Tel. (+31-70) 340-2040, Tx. 31 651 epo nl.
Fax: (+31-70) 340-3016

Authorized officer

Bork, A-M

C.(Continuation) DOCUMENTS CONSIDERED TO BE RELEVANT		
Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	EP 0 560 128 A (BASF AKTIENGESELLSCHAFT) 15 September 1993 (1993-09-15) cited in the application column 2, line 44 - column 3, line 9 column 3, line 46 - column 7, line 8 column 8, line 1 - line 16 column 8, line 47 - column 10, line 26; example 1 -----	1-4, 9-11
X	US 4 035 560 A (CAUMARTIN ET AL) 12 July 1977 (1977-07-12) column 1, line 54 - column 2, line 60; claims 1,6-9; examples 1-6 -----	1,3,4,9, 10
X	US 5 661 098 A (HARRISON ET AL) 26 August 1997 (1997-08-26) column 2, line 60 - line 67 column 3, line 22 - column 4, line 43; example 1 -----	1-4
X	US 5 446 001 A (GURTZGEN ET AL) 29 August 1995 (1995-08-29) column 2, line 29 - line 37 column 3, line 65 - column 4, line 44; claims 12,13,19,20; examples 1,2 -----	1-4
A	PATENT ABSTRACTS OF JAPAN vol. 002, no. 052 (C-011), 14 April 1978 (1978-04-14) & JP 53 011982 A (MITSUBISHI PETROCHEM CO LTD), 2 February 1978 (1978-02-02) cited in the application abstract -----	1-4, 9

INTERNATIONAL SEARCH REPORTInternational application No.
PCT/EP2004/011316**Box II Observations where certain claims were found unsearchable (Continuation of item 2 of first sheet)**

This International Search Report has not been established in respect of certain claims under Article 17(2)(a) for the following reasons:

1. Claims Nos.: because they relate to subject matter not required to be searched by this Authority, namely:

2. Claims Nos.: **partially 1-11** because they relate to parts of the International Application that do not comply with the prescribed requirements to such an extent that no meaningful International Search can be carried out, specifically:
see FURTHER INFORMATION sheet PCT/ISA/210

3. Claims Nos.: because they are dependent claims and are not drafted in accordance with the second and third sentences of Rule 6.4(a).

Box III Observations where unity of invention is lacking (Continuation of item 3 of first sheet)

This International Searching Authority found multiple inventions in this International application, as follows:

1. As all required additional search fees were timely paid by the applicant, this International Search Report covers all searchable claims.

2. As all searchable claims could be searched without effort justifying an additional fee, this Authority did not invite payment of any additional fee.

3. As only some of the required additional search fees were timely paid by the applicant, this International Search Report covers only those claims for which fees were paid, specifically claims Nos.:

4. No required additional search fees were timely paid by the applicant. Consequently, this International Search Report is restricted to the invention first mentioned in the claims; it is covered by claims Nos.:

Remark on Protest

The additional search fees were accompanied by the applicant's protest.
 No protest accompanied the payment of additional search fees.

FURTHER INFORMATION CONTINUED FROM PCT/ISA/ 210

Continuation of Box II.2

Claims Nos.: partially 1-11

Present independent claim 1 relate to a product defined (inter alia) by reference to the following parameter(s):

P1: angle of repose

The use of this parameter in the present context is considered to lead to a lack of clarity within the meaning of Article 6 PCT. It is impossible to compare the parameters the applicant has chosen to employ with what is set out in the prior art. The lack of clarity is such as to render a meaningful complete search impossible. Consequently, the search has been restricted to: pulverulent solids according to claim 1 consisting essentially of at least one metal alkyl compound bound to a finely divided, porous, mechanically stable and chemically inert support, which have a proportion by weight of metal alkyl compound of at least 5% by weight, based on the support, their preparation and use.

The applicant's attention is drawn to the fact that claims relating to inventions in respect of which no international search report has been established need not be the subject of an international preliminary examination (Rule 66.1(e) PCT). The applicant is advised that the EPO policy when acting as an International Preliminary Examining Authority is normally not to carry out a preliminary examination on matter which has not been searched. This is the case irrespective of whether or not the claims are amended following receipt of the search report or during any Chapter II procedure. If the application proceeds into the regional phase before the EPO, the applicant is reminded that a search may be carried out during examination before the EPO (see EPO Guideline C-VI, 8.5), should the problems which led to the Article 17(2) declaration be overcome.

INTERNATIONAL SEARCH REPORT

Information on patent family members

International Application No
PCT/EP2004/011316

Patent document cited in search report	Publication date		Patent family member(s)	Publication date
WO 9510542	A	20-04-1995	US 5466649 A DE 69421287 D1 DE 69421287 T2 EP 0723557 A1 ES 2139758 T3 WO 9510542 A1 US 5516737 A US 5516861 A	14-11-1995 25-11-1999 20-07-2000 31-07-1996 16-02-2000 20-04-1995 14-05-1996 14-05-1996
US 4370456	A	25-01-1983	AU 548970 B2 AU 8893782 A CA 1187069 A1 DE 3266704 D1 EP 0080052 A1 GR 76707 A1 JP 63033404 A JP 1432671 C JP 58132012 A JP 62041602 B ZA 8207277 A	09-01-1986 02-06-1983 14-05-1985 07-11-1985 01-06-1983 29-08-1984 13-02-1988 24-03-1988 06-08-1983 03-09-1987 31-08-1983
SU 409723	A	05-01-1974	SU 409723 A1	05-01-1974
EP 0560128	A	15-09-1993	DE 4207899 A1 DE 59302413 D1 EP 0560128 A2 ES 2086146 T3 JP 3585508 B2 JP 6023273 A	16-09-1993 05-06-1996 15-09-1993 16-06-1996 04-11-2004 01-02-1994
US 4035560	A	12-07-1977	FR 2312512 A1 BE 842241 A1 DE 2623693 A1 ES 448203 A1 GB 1500868 A IT 1071479 B LU 75016 A1 NL 7605598 A ,B,	24-12-1976 26-11-1976 02-12-1976 01-07-1977 15-02-1978 10-04-1985 20-01-1977 30-11-1976
US 5661098	A	26-08-1997	CA 2163681 A1	25-05-1997
US 5446001	A	29-08-1995	DE 4336659 A1 CA 2127671 A1 DE 59407302 D1 EP 0650967 A1 FI 945043 A JP 7188253 A NO 944054 A	04-05-1995 28-04-1995 24-12-1998 03-05-1995 28-04-1995 25-07-1995 28-04-1995
JP 53011982	A	02-02-1978	NONE	